

<b>Module Title:</b>	Chemistry 1
<b>Language of Instruction:</b>	English
<b>Credits:</b>	5
<b>NFQ Level:</b>	6
<b>Module Delivered In</b>	<a href="#">6 programme(s)</a>
<b>Teaching &amp; Learning Strategies:</b>	This module will be taught in two theory classes of 1 hour duration plus one 3 hour practical per week. To consolidate lectures and practicals, students will be required to carry out assignments and prepare a weekly practical report analysing their own research and results. Any course –related issue of questions that may arise will be discussed at lectures.
<b>Module Aim:</b>	To provide students with a knowledge of basic theoretical and practical chemistry principles.
<b>Learning Outcomes</b>	
<i>On successful completion of this module the learner should be able to:</i>	
LO1	Describe basic chemical terminology, facts and principles.
LO2	Compute stoichiometric, titrimetric, and other relevant calculations.
LO3	Perform practical laboratory skills in chemical and analytical procedures
<b>Pre-requisite learning</b>	
<b>Module Recommendations</b>	
<i>This is prior learning (or a practical skill) that is recommended before enrolment in this module.</i>	
No recommendations listed	
<b>Incompatible Modules</b>	
<i>These are modules which have learning outcomes that are too similar to the learning outcomes of this module.</i>	
No incompatible modules listed	
<b>Co-requisite Modules</b>	
No Co-requisite modules listed	
<b>Requirements</b>	
<i>This is prior learning (or a practical skill) that is mandatory before enrolment in this module is allowed.</i>	
No requirements listed	

**Module Content & Assessment**

Indicative Content
<b>Atomic Theory</b> History & modern view of atomic structure. Electronic configuration, Pauli's exclusion principle, the Aufbau principle, and Hund's rule. Isotopes. Atomic Mass.
<b>Periodic Table</b> Periodic table. How to read the table. Periodicity and trends, to include atomic radii, ionization energy, electron affinity, and electronegativity.
<b>Bonding and Intermolecular Interactions</b> Molecules/Ions/Metallic solids. Chemical bonds. Polarity. Van Der Waals Forces (London dispersion, dipole-dipole). Hydrogen bonding. States of matter - characteristics of solids, liquids, gases. Density.
<b>Stoichiometry</b> Calculations with chemical formulae and equations: Formulae. Elemental analysis. Molecular mass. The mole. Chemical equations and stoichiometry.
<b>Volumetric Analysis</b> Solutions. Concentrations: Molarity, ppm, percent solutions. Dilutions. Titrations: Apparatus, safety procedures, and technique. Titrimetric calculations.
<b>Chemical Equilibrium</b> Dynamic equilibrium. Le Chatelier's principle to demonstrate effects on equilibrium. Applications in control of industrial processes. Haber process.
<b>Acid base equilibria</b> Acids and bases - Bronsted-Lowry acids and bases. pH scale. Strong and weak acids. Buffer solutions.
<b>Practicals</b> Use of volumetric glassware. Titrimetric analysis including direct, indirect, and back titrations using various acid-base examples. pH determination and equilibria reactions. Qualitative inorganic analysis.

Assessment Breakdown	%
Continuous Assessment	20.00%
Practical	50.00%
End of Module Formal Examination	30.00%

**Special Regulation**

Students must achieve a minimum grade (35%) in both the practical/CA and final examination

**Continuous Assessment**

Assessment Type	Assessment Description	Outcome addressed	% of total	Assessment Date
Examination	Throughout Year	1,2	20.00	n/a

No Project

**Practical**

Assessment Type	Assessment Description	Outcome addressed	% of total	Assessment Date
Practical/Skills Evaluation	Practical Log Book	3	50.00	Every Week

**End of Module Formal Examination**

Assessment Type	Assessment Description	Outcome addressed	% of total	Assessment Date
Formal Exam	3 hour exam	1,2	30.00	End-of-Semester

SETU Carlow Campus reserves the right to alter the nature and timings of assessment

**Module Workload**

<b>Workload: Full Time</b>		
<i>Workload Type</i>	<i>Frequency</i>	<i>Average Weekly Learner Workload</i>
Lecture	12 Weeks per Stage	2.00
Practicals	12 Weeks per Stage	3.00
Independent Learning	15 Weeks per Stage	4.33
Total Hours		125.00

**Module Delivered In**

Programme Code	Programme	Semester	Delivery
CW_SABTP_B	<a href="#">Bachelor of Science (Honours) in Biosciences with Biopharmaceuticals</a>	1	Mandatory
CW_SABRE_B	<a href="#">Bachelor of Science (Honours) in Brewing and Distilling</a>	1	Mandatory
CW_SAPHA_B	<a href="#">Bachelor of Science (Honours) in Pharmaceutics and Drug Formulation</a>	1	Mandatory
CW_SAASC_D	<a href="#">Bachelor of Science in Analytical Science</a>	1	Mandatory
CW_SABFQ_D	<a href="#">Bachelor of Science in Biosciences</a>	1	Mandatory
CW_SASCI_C	<a href="#">Higher Certificate in Science in Applied Biology or Applied Chemistry</a>	1	Mandatory